

Answers

Chemistry 11

Unit 2 test review

TOOMBS

Practicing names and formulas:

SF!
UNITS!!!
⇓

Ion Combination	I or C	Formula	Name	Molar Mass	g/mole
$\text{Ca}^{+2} / \text{NO}_2^-$	I	$\text{Ca}(\text{NO}_2)_2$	calcium nitrite	132.1	
$\text{B}^{+3} / \text{NO}_4^-$	I	$\text{B}(\text{NO}_4)_3$	boron pernitrate	244.8	
$\text{Si}^{+4} / \text{Cl}^-$	C	SiCl_4	silicon tetrachloride	170.1	
$\text{NH}_4^+ / \text{Cl}^-$	I	NH_4Cl	ammonium chloride	53.6	
	I	Al_2S_3	aluminum sulphide	150.3	
	I	FeSO_4	iron (II) sulphate	151.9	
	C	SF_6	sulphur hexafluoride	146.1	
	I	PtF_6	platinum (VI) fluoride	309.1	
	C	Cl_2O_7	dichlorine heptoxide	183.0	
	C	PO_5	phosphorus pentoxide	111.0	
	I	H_3PO_5	perphosphoric acid	114.0	
	I	$\text{Ca}(\text{CH}_3\text{COO})_2$	Calcium acetate	158.2	
	C	N_2O_5	Dinitrogen pentoxide	108.0	
	I	$\text{Mn}(\text{OH})_2$	Manganese (II) hydroxide	88.9	
	I	Cr_2O_3	Chromium (III) oxide	152.0	
	C	N_2Br_3	Nitrogen tribromide	253.7	
	I	$\text{Mg}_3(\text{AsO}_3)_2$	Magnesium arsenite	318.7	
More or acid molecules to come	I	HCl	hydrochloric acid	36.5	
	I	H_2CO_3	Carbonic acid	62.0	
	I	CH_3COOH	acetic acid	60.0	
	I	H_3PO_4	Phosphoric acid	98.0	
	I	H_2S	hydro sulphuric acid	34.1	
	I	H_2SO_4	sulphuric acid	98.1	
	I	H_3BO_3	Boric acid	61.8	
	I	HNO_2	nitrous acid	47.0	

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