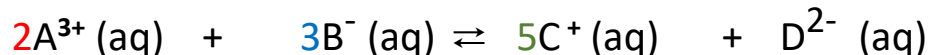


I. R. E. ANALYSIS

1. 12 M of A^{3+} (aq) are reacted with 14 M of B^{-} (aq) as shown in the following equation. At equilibrium, 20 M of C^{+} (aq) are measured.

How much of the A^{3+} (aq) and B^{-} (aq) reacted? How much of the A^{3+} (aq) and B^{-} (aq) and D^{2-} (aq) are present at equilibrium?



INITIAL	12 M	14 M	-----	-----
REACTION				
EQUILIBRIUM			20 M	

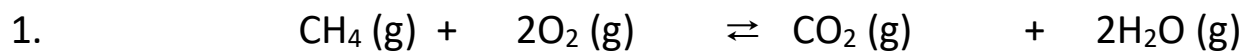
2. $X (g) + 3Y (g) \rightleftharpoons 6Z (g)$

I	1.6 M	2.4 M	2.1 M
R			
E		3.0 M	

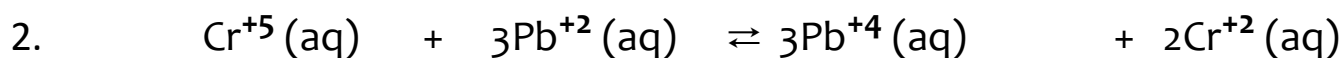
3. $3H_2 (g) + N_2 (g) \rightleftharpoons 2NH_3 (g)$

I	5.0 M	3.0 M	2.0 M
C			
E			5.0 M

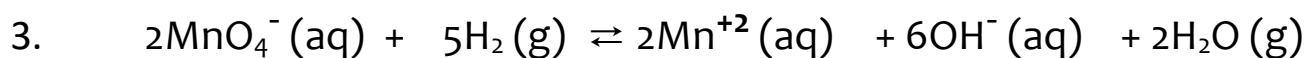
2.1 ESSENTIAL EQUILIBRIUM EXERCISE



I	8.0 M	11.0 M	-----	-----
C				
E			3.0 M	



I	0.84 M	0.75 M	-----	-----
R				
E	0.60 M			



I	5.50 M	7.00 M	4.00 M	1.50 M	6.00 M
C					
E				0.54 M	

