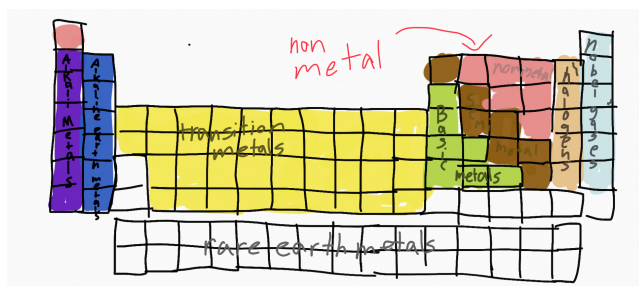


Review #1-16

1. A. Because it is not creating a new substance.
2. D because when a electron is outside the nucleus it is 1/1800
3. C
4. Metals have lustre, are normally malleable and ductile and conduct heat
Non metals aren't lustrous, are brittle, don't conduct heat
Metalloids have some of both characteristics
- 5.



6. a) H b) He c) Na d) B e) N f) S g) O

7. a) Oxygen b) aluminum c) Helium d) Potassium

8 Halogens: highly reactive, bright colours as gasses, gains 1 electron from ion

Noble gasses: unreactive, colourless as gasses, do not readily form ions

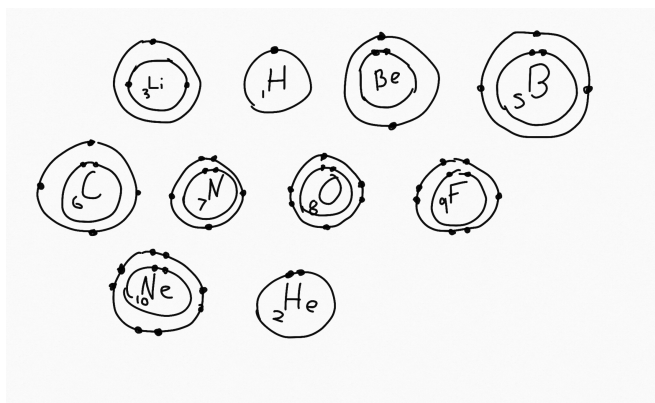
Alkaline earth metals: hard metal, low density, loses two electrons to form ion

Alkali metals: soft metals, react with water to form hydrogen, loses 1 electron from ion

9. A. 1

10. The electrons and protons are both the atomic number the neutron is the mass - the atomic number.

11.



12. C.

13. a) 1: 14 2: 10 3: 8

b) 1. silicon 2. Neon 3. oxygen

c) 14, 10, 8

d) 14, 10, 6

e) 14, 10, 8

f) 28, 20, 14