Date: $\qquad$

Given the following information about the "charges" or "values" that certain elements or groups of elements have:

|  | $\mathbf{+ 1}$ | $\mathbf{+ 2}$ | $\mathbf{+ 3}$ | ??? | $\mathbf{- 1}$ | $\mathbf{- 2}$ | $\mathbf{- 3}$ |
| :--- | :--- | :--- | :--- | :--- | :--- | :--- | :--- |
| ELEMENTS | H | Mg | B | Sn | F | O | N |
| Or groups of <br> elements | Li | Ca | Al | Cu | Cl | S | P |
| that had | Na | Sr | Ge | Fe | Br | $\mathrm{SO}_{4}$ | $\mathrm{BO}_{3}$ |
| this identity <br> Or "charge" | K |  |  | Ni | $\mathrm{NO}_{4}$ |  |  |
| Au | $\mathrm{CO}_{3}$ | $\mathrm{PO}_{4}$ |  |  |  |  |  |

For the following given molecules, cross out the incorrect formulas.
$\qquad$

AIP
$\mathrm{K}_{2} \mathrm{O}$
$\mathrm{Au}_{2} \mathrm{O}_{3}$
$\mathrm{AuPO}_{4}$

HCl

NP

MgO
$\mathrm{NiBr}_{3}$
$\mathrm{SnCl}_{2}$
$\mathrm{H}_{3} \mathrm{BO}_{3}$

AIB
$\mathrm{NH}_{4} \mathrm{Cl}$
$\mathrm{CaCO}_{3}$
$\mathrm{SnCl}_{4}$
$\mathrm{H}_{2} \mathrm{CO}_{3}$

In the space below, write as many possible Formulas for Chemical Compounds (or Molecules), as you can - do NOT repeat any of the examples that were used above.

